1605, 1416, 1376, 1355, 1326, 1290, 1171, 1125, 1107, 1091, 1067, 1007, 978, 930, 913, 778, 756, 740, 729 cm^{-1}. ^1H NMR ($CDCl_3$): δ = 3.53 (d, 2 H, tz—H^a), 7.34 (d, 2 H, tz—H^b), 7.93 (m, 8 H, Pc—H^2), 9.15 (m, 8 H, Pc—H^1). ^{13}C NMR ($CDCl_3$): δ = 122.7 (C_1), 129.9 (C_2), 139.0 (C_{4a}), 145.6 (N=C—N), 157.3 (tz—C_a), 160.3 (tz—C_b). UV/vis ($CHCl_3$): λ = 618, 570 (sh), 365, 305 nm. Anal. Calc for $C_{36}H_{20}N_{16}Os$ (866): C, 49.88; H, 2.31; N, 25.87. Found: C, 50.28; H, 2.63; N, 20.43.

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Additions and Corrections

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Janet R. Morrow and William C. Trogler*: Hydrolysis of Phosphate Diesters with Copper(II) Catalysts.

Page 3390. Table III of the original paper suffered from an erroneous data entry in the program used to calculate the species present at equilibrium. The equilibrium constants used and the corrected table follow. The new values suggest $Cu(OH)(bpy)^+$ is the major species in solution with a 1:1 Cu:bpy ratio, not the dimer. This removes the inconsistency between the kinetics data, which showed a simple first-order dependence in $[Cu^{2+}]_{tot}$, and the previous calculation of equilibrium species, which predicted significant dimer formation. This change strengthens the conclusions in the paper.

$$[Hbpy^+]/[bpy][H^+] = 1.13 \times 10^4$$

$$[Cu(bpy)^{2+}]/[Cu^{2+}][bpy] = 6.47 \times 10^6$$

$$[Cu(bpy)_2{}^{2+}]/[Cu^{2+}][bpy]^2 = 5.15 \times 10^{11}$$

$$[Cu(bpy)(OH)^+]/[Cu(bpy)(OH)_2][H^+] = 4.06 \times 10^{10}$$

$$[Cu(bpy)^{2+}]/[H^+][Cu(bpy)(OH)^+] = 1.15 \times 10^7$$

$$[(Cu(bpy)(OH))_2{}^{2+}]/[Cu(bpy)(OH)^+]^2 = 8.65 \times 10^3$$

Table III. Effect of the Concentration (M) of 2,2′-Bipyridine on the Concentrations (M) of Copper(II) Complexes in Water at 75 °C[a]

[bpy]	$[Cu^{2+}]$	$[Cu(bpy)^{2+}]$	$[Cu(bpy)_2{}^{2+}]$	$[Cu(bpy)(OH)^+]$	$[(Cu(bpy)(OH))_2{}^{2+}]$
2.5×10^{-5}	2.5×10^{-5}	2.8×10^{-6}	3.8×10^{-9}	1.7×10^{-5}	2.5×10^{-6}
5.0×10^{-5}	1.0×10^{-6}	4.7×10^{-6}	2.7×10^{-7}	2.9×10^{-5}	7.4×10^{-6}
1.0×10^{-4}	1.6×10^{-8}	3.9×10^{-6}	1.2×10^{-5}	2.4×10^{-5}	5.0×10^{-6}
2.0×10^{-4}	3.3×10^{-9}	2.6×10^{-6}	2.6×10^{-5}	1.6×10^{-5}	2.3×10^{-6}

[a] bpy = 2,2′-bipyridine; pH 7.85; $[Cu^{2+}] = 5.0 \times 10^{-5}$ M. The $[Cu(bpy)(OH)_2]$ calculated was below 10^{-7} M in all cases.

—William C. Trogler